










Sample Literacy Strategy for Scientific Reading Passages

Glycine, $\text{NH}_2\text{CH}_2\text{COOH}$, is an organic compound found in proteins. Acetamide, CH_3CONH_2 , is an organic compound that is an excellent solvent. Both glycine and acetamide consist of the same four elements, but the compounds have different functional groups.

Original Regents Question: Draw the structural formula for acetamide.

Where to look?	My Question/Your Answer
	What is the formula for acetamide?
	What functional groups are found in acetamide?
	Use your reference table to draw what these functional groups look like.
	What is the difference between the formula that is given for acetamide and a structural formula?
	How many bonds must carbon form to be stable?
	How many bonds must nitrogen form to be stable?
	How many bonds must hydrogen form to be stable?
	How many bonds must oxygen form to be stable?
	Draw a structural formula for acetamide.

